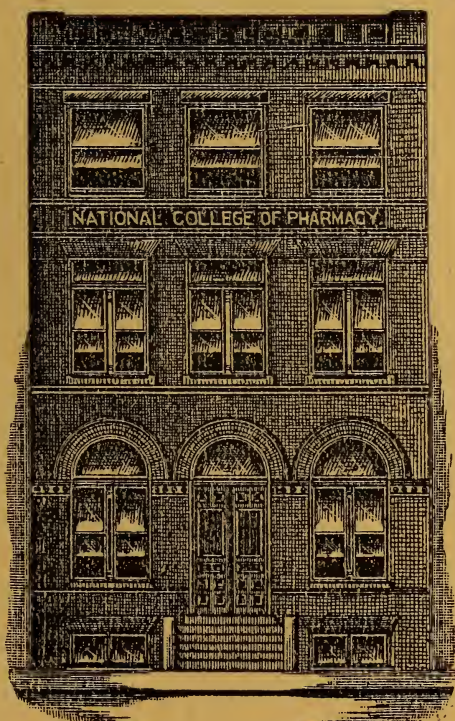


G2932^Cph
1911/12

The George Washington University National College of Pharmacy

Announcement for 1911-1912

THE LIBRARY
OF THE
UNIVERSITY OF MICHIGAN



Fortieth Annual Session

JUN 16 1912

808 J Street N. W.

Washington, D. C.

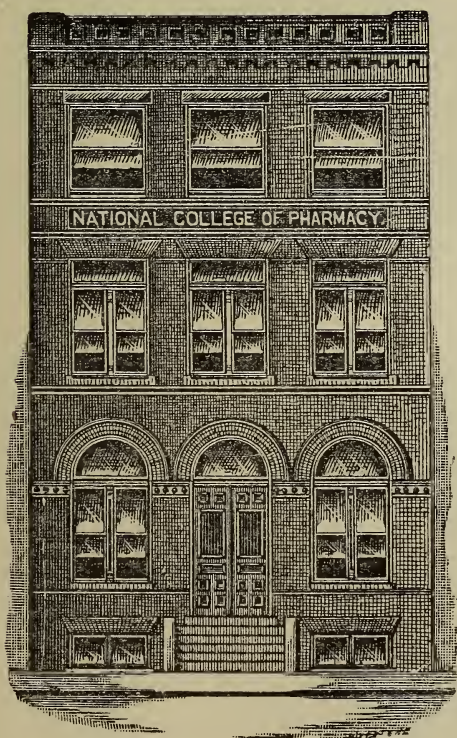
COLLEGE CALENDAR.—The Fortieth Annual Session of the College will begin on Wednesday, September 20, 1911, and close on Wednesday, June 5, 1912.

MONDAY.	TUESDAY.	WEDNESDAY.	THURSDAY.	FRIDAY.	SATURDAY.
FRESHMEN.		JUNIORS.		SENIORS.	
Botany. Lectures & Recitations. 10-11 A. M. Pharmacy. Lectures & Recitations. 11 A. M.-12 M. Recess, 12 M.-12:30 P. M. Pharmacy. Laboratory Work. 12:30-4 P. M.	Botany and Materia Medica. Lectures & Recitations. 10-11 A. M. Pharmacy. Lectures & Recitations. 11 A. M.-12 M. Recess, 12 M.-12:30 P. M. Pharmacy. Laboratory Work. 12:30-4 P. M.	Botany and Materia Medica. Lectures & Recitations. 10-11 A. M. Pharmacy. Lectures & Recitations. 11 A. M.-12 M. Recess, 12 M.-12:30 P. M. Pharmacy. Laboratory Work. 12:30-4 P. M.		Materia Medica and Toxicology. Lectures & Recitations. 10-11 A. M. Pharmacy. Lectures & Recitations. 11 A. M.-12 M. Recess, 12 M.-12:30 P. M. Pharmacy. Laboratory Work. 12:30-4 P. M.	
MONDAY.	TUESDAY.	WEDNESDAY.	THURSDAY.	FRIDAY.	SATURDAY.
JUNIORS.	SENIORS.	FRESHMEN.	SENIORS.	JUNIORS.	
Physics and General Chemistry. Lectures & Recitations. 6-7 P. M. Analytical Chemistry Lectures, Recitations and Laboratory Work. 7-11 P. M.	Microscopy. Lectures and Practice. 6-8 P. M. Mercantile Pharmacy Oct. 3-Jan. 23 Lectures and Practice 8-10 P. M. Pharmaceutical Jurisprudence Jan. 30 to End of Term. 8-9 P. M.	Physics and General Chemistry. Lectures & Recitations. 6-7 P. M. Analytical Chemistry Lectures, Recitations and Laboratory Work. 7-11 P. M.	General and Organic Chemistry. Lectures & Recitations. 6-7 P. M. Quantitative Chemical Analysis Lectures, Laboratory Work and Recitations. 7-11 P. M.	Microscopy. Lectures, Laboratory Work and Recitations. 6-8 P. M.	

Entrance examination at 1 p. m. on Thursday, September 14, 1911, in the Lecture-room of the College. Annual examination of Freshmen and Juniors for promotion and of Seniors for graduation begins on Friday, April 26, 1912.
 1911—November 30, legal holiday; no College exercises. December 21, last lecture before Christmas vacation. 1912—January 3, lectures resume.
 February 22, Legal Holiday June 5, graduation day; College closes.

The George Washington University National College of Pharmacy

Announcement for 1911-1912



Fortieth Annual Session

808 J Street N. W.

Washington, D. C.

MEMBERS OF THE COLLEGE.

PHILIP J. AFFLECK
 THOMAS H. ATKINSON
 HERBERT N. BEALL.
 GEORGE W. BOYD.
 W. EDWARD BOYD.
 WILLIAM D. BRACE.
 WYMOND H. BRADBURY.
 ALFRED T. BRONAUGH.
 HOWARD M. BRADBURY.
 ROBERT F. BOGGAN
 LOUIS F. BRADLEY
 CHARLES B. CAMPBELL.
 EUGENE E. CISSEL.
 ALBERT N. CONNER.
 FRANCIS M. CRISWELL.
 PRESTON C. DAY.
 WILLIAM C. DOWNEY.
 PETER J. DUNCAN
 ROGER W. DUFFEY
 HERBERT C. EASTERDAY.
 JAMES K. EPPLEY.
 VICTOR H. ESCH.
 HENRY EVANS.
 W. ASHTON EVANS
 MARTIN S. FEALY.
 LEWIS FLEMER.
 LEOPOLD H. FOSTER.
 JOSEPH D. FRANZONI.
 CHARLES J. FUHRMANN.
 MALCOLM G. GIBBS.
 EDWARD GREEN.
 CHARLES E. GROSS.
 WILLIAM G. GENTNER
 J. MURRAY HACKETT.
 ROBERT N. HARPER.
 EDWARD A. HELMSEN.
 RUFUS K. HELPHENSTINE.
 FRANK C. HENRY.
 WILLIAM P. HERBST.
 SAMUEL L. HILTON.
 R. CLIFFORD HINES.
 JAMES T. HOSKINS.
 FLORENCE V. HOSKINS
 CHARLES HAWKINS.
 R. VERNON HOUSTON
 GEORGE W. HURLEBAUS.
 FRED F. HAFELFINGER
 WALTER R. HILL

HENRY A. JOHNSTON.
 JOHN R. JACOBS
 T. A. T. JUDD.
 WILLIAM P. KENEALY
 HENRY E. KALUSOWSKI.
 LYMAN F. KEBLER.
 WILLIAM T. KERFOOT.
 CHARLES J. LENNON.
 WILLIAM H. McCLURE
 JOHN R. MAJOR.
 GEO. T. MANKIN.
 J. WILLARD McCHESNEY.
 M. H. MENKE.
 T. K. MELSON.
 WILLIAM F. MATTINGLY
 GUY M. NEELY.
 THOMAS E. OGRAM.
 OSCAR OLDBERG (Honorary).
 PAUL PEARSON.
 FRANK PITZER.
 R. LUCIEN QUIGLEY.
 ALBERT M. READ (Honorary)
 CHARLES D. REMSBURG.
 FRANK R. RICHARDSON.
 W. S. RICHARDSON.
 EARLE K. RICHARDSON
 SAMUEL A. RICHARDSON
 J. HARRY STUTT
 FRED SCHAFHIRT.
 J. FRENCH SIMPSON.
 HARRY C. SNYDER.
 H. E. SPRUCEBANK.
 FRANK T. STONE.
 SAMUEL T. STOTT.
 ODEN R. SUDLER.
 FRANK B. TIPTON
 AUGUSTUS C. TAYLOR
 F. A. TSCHIFFELY.
 ROBERT A. VEITCH.
 MARTIN I. WILBERT
 R. E. LEE WILLIAMSON
 SAMUEL WAGGAMAN
 S. M. WAGNER.
 FRANK P. WELLER.
 EDWARD W. WHITESIDE
 CHARLES S. WALTER
 CONRAD H. WEISS

OFFICERS AND COMMITTEES
OF
NATIONAL COLLEGE OF PHARMACY.

SESSION OF 1911-1912.



President.

REAR ADMIRAL CHARLES H. STOCKTON, LL. D.

Dean and Chairman.

HENRY E. KALUSOWSKI.

Vice-Chairman.

LEWIS FLEMER

Secretary.

WYMOND H. BRADBURY.

Treasurer.

H. C. EASTERDAY.

TRUSTEES,

With date of expiration of term of office.

CHARLES B. CAMPBELL, 1914.

HERBERT C. EASTERDAY, 1914.

LEWIS FLEMER, 1914.

WILLARD S. RICHARDSON, 1914.

FRANK C. HENRY, 1912.

SAMUEL L. HILTON, 1912.

HENRY E. KALUSOWSKI, 1912.

AUGUSTUS C. TAYLOR, 1912.

WYMOND H. BRADBURY, 1913.

JAMES K. EPPLEY, 1913.

SAMUEL WAGGAMAN, 1913.

FRANK P. WELLER, 1913.

STANDING COMMITTEES.

Finance and Business.

WILLARD S. RICHARDSON.

JAMES K. EPPLEY.

CHARLES B. CAMPBELL.

Pharmaceutical Education and Conduct of School of Pharmacy.

HENRY E. KALUSOWSKI. FRANK C. HENRY. SAMUEL L. HILTON.

LEWIS FLEMER.

FRANK P. WELLER.

U. S. Pharmacopœia and Progress of Pharmacy.

SAMUEL L. HILTON.

H. E. KALUSOWSKI.

LEWIS FLEMER.

National Formulary.

MARTIN I. WILBERT. AUGUSTUS C. TAYLOR. GEO. W. HURLEBAUS.

Publications, Library and Cabinet.

SAMUEL WAGGAMAN.

CHAS. HAWKINS.

CHARLES E. GROSS

FACULTY

HENRY E. KALUSOWSKI, M. D., PHAR. D., DEAN,
Professor of Pharmacy.

SAMUEL WAGGAMAN, M. D., PHAR. D.,
Professor of Materia Medica, Botany and Toxicology

GEORGE A. MENGE, PH. D.,
Professor of Chemistry and Physics.

HOWARD M. BRADBURY, PHAR. D.,
Professor of Analytical Chemistry.

BURTON J. HOWARD, B. S.,
Professor of Microscopy.

ALEXANDER MUNCASTER, PHAR. D., LL. B., LL. M.,
Professor of Pharmaceutical Jurisprudence

HENRY B. FLOYD, PHAR. D.,
Professor of Mercantile Pharmacy.

DOUGLAS TSCHIFFELY, PHAR. D.,
Assistant to the Professor of Pharmacy

HOMER K. BUTLER, PHAR. D.,
LORING W. BEESON, PHAR. D.,
Assistants to the Professors of Chemistry.



BOARD OF EXAMINERS

HENRY E. KALUSOWSKI, CHAIRMAN.

LEWIS FLEMER.

FRANK C. HENRY.

SAMUEL L. HILTON.

BURTON J. HOWARD.

ALEXANDER MUNCASTER.

SAMUEL WAGGAMAN.

H. M. BRADBURY.

F. P. WELLER.

GEORGE A. MENGE.

HENRY B. FLOYD.

THE GEORGE WASHINGTON UNIVERSITY NATIONAL COLLEGE OF PHARMACY

No. 40

WASHINGTON, D. C.

1911-1912.

THE NATIONAL COLLEGE OF PHARMACY, of Washington, D. C., was organized and began teaching November 11, 1872.

In February, 1906, it became a member of the educational system of The George Washington University under the charter of the University granted by Congress March 3, 1905, providing for the organization of colleges. The President of the University is *ex-officio* President of the National College of Pharmacy, and the College is represented in the President's Council by its Dean.

The official title of the College now is THE GEORGE WASHINGTON UNIVERSITY NATIONAL COLLEGE OF PHARMACY.

COLLEGE BUILDING.

The College building is centrally located on I Street, Northwest, between 8th and 9th Streets. It is easily reached by any street car line of the city.

LOCAL ADVANTAGES.

Washington is the most attractive and beautiful city in the country, as well as one of the most healthful. Its extensive Libraries, Museums, and Botanical Gardens are open to Students every day except Sunday. The facilities thus afforded for studying Botany and Materia Medica are unequaled anywhere.

PRELIMINARY EDUCATION AND EXAMINATION

The college requires of its Matriculants a knowledge of the branches usually taught in the public schools of Washington, D. C. to the extent of *one year in the high school or its equivalent*, and they shall be at least seventeen years of age. *beginning with the session of 1912-13, candidates for admission to the college will be required to have two years High School instruction or the equivalent; beginning with the session of 1914-15, three years High School instruction or the equivalent; beginning with the session of 1916-17 four years High School or the equivalent.* Evidence of this may be shown by certificates from reputable teachers or by the results of an examination to be held at the College, at one o'clock p. m., THURSDAY, September 14, 1911.

THREE YEARS' STUDY REQUIRED BEFORE GRADUATION.

The rapid progress made in the sciences and arts directly affecting to practice of Pharmacy, and the widely diversified knowledge now required to keep the pharmacist in touch with the best results of modern research, have made an extension of the course of instruction necessary. Students entering the College will therefore be required to take a course of study covering a period of three years before they will be entitled to graduation. Upon entry students will be assigned to the first or Freshman Class, from which, after passing a satisfactory examination, they will be promoted to the Junior and Senior Classes at the end of the first and second scholastic years respectively.

MATERIA MEDICA, BOTANY AND TOXICOLOGY

PROF. SAMUEL WAGGAMAN

FRESHMAN COURSE

The first four or five lectures will constitute an introduction to the study of Elementary Botany, after which will follow in order Vegetable Histology and Plant Physiology. The subject-matter of these lectures will be thoroughly explained by means of charts, diagrams, specimens and illustrations.

JUNIOR COURSE

The lectures in this course will be devoted to the consideration of the various theories concerning the vegetable world and the practical results obtained by experienced laborers in this vast field of science. During the time devoted to these subjects the lecturer will illustrate the lectures by means of the lantern and microscope, exhibiting such parts of life as are known to be of practical importance. A large part of the course will be devoted to a consideration of the official organic drugs.

SENIOR COURSE

The lectures following the above-mentioned courses on plants and their relation to Materia Medica will be mainly upon the active principles, adulterants, official preparations, therapeutic uses, and doses; after which the organic and inorganic poisons will be taken up under three heads: 1st, Corrosive; 2d, Irritant; and 3d, Neurotic Poisons. Under these three divisions will be explained briefly their action, detection and antidote.



PHARMACY

PROF. HENRY E. KALUSOWSKI

DOUGLAS TSCHIFFELY

Assistant

This course will be devoted to the study of the various pharmaceutical processes and operations. The opening lecture will define Pharmacy and state its relations to the arts and sciences; then will follow in the order named lectures on Metrology, Heat, Thermometry, Evaporation, Distillation, Fusion, Sublimation, Calcination, Granulation, Comminution, Solution, Filtration, Clarification, Decoloration, Precipitation, Crystallization and Extraction, during which the various methods used to bring about the desired results will be illustrated.

JUNIOR COURSE.

A part of this course will be taken up with the study of official preparations obtained from the mineral kingdom, beginning with Bromine, Chlorine, Iodine, Phosphorus, and Sulphur, followed in the order named by Carbon, Boron, Silicon, the inorganic acids, Potassium, Sodium, Lithium, Ammonium, Magnesium, Calcium, Barium, Zinc, Aluminum, Cerium, Cadmium, Manganese, Iron, Chromium, Lead, Silver, Copper, Mercury, Antimony, Arsenic, Bismuth, and Gold, followed by a study of Fixed Oils and Fat, Volatile Oils, Alkaloids, their sources and separation, and animal products.

SENIOR COURSE.

The time during this course will be mainly given to the study of compounds chiefly derived from organic matter, such as Cellulose and products obtained therefrom, Amylaceous, Mucilaginous and Saccharine substances; Soaps, Resinoids and products from resins, to be followed by a systematic instruction in compounding and dispensing prescriptions.

In addition to oral quizzes each student in all classes will receive at regular periods a series of written questions covering the subject-matter of previous lectures, answers to which will have to be returned on the day of issuing.



ANALYTICAL CHEMISTRY.

PROF. HOWARD M. BRADBURY.

HOMER K. BUTLER, Assistant.

The instruction in this department is intended to present to the student the chemical tests of the United States Pharmacopœia; to familiarize him with methods for the identification of substances and for the detection of impurities; to instruct him in the methods of assaying and the use of volumetric solutions, and to enable him to analyze any ordinary mixture of inorganic material.

The course of instruction embraces three years of practice in the chemical laboratory and class-room exercises.

For the purpose of carrying out the work of this department a large, well-equipped laboratory is provided having drawer and locker accommodations for one hundred and twenty students and desk space for forty students working together at one time. The laboratory is provided with the usual water and gas facilities, and has been wired and installed with electric apparatus whereby electro-chemical methods of analysis can be taught and the application of the electric current to the preparation of chemicals by the methods of electro-chemistry can be illustrated before the students. The laboratory is also provided with ample hood facilities for keeping the atmosphere of the room free as possible from deleterious fumes, and is equipped with means for giving an ample supply of distilled water. These and other facilities afford students exceptional opportunities to become familiar with the fundamental principles of the science of chemistry.

The first year is devoted to experimental work so arranged as to supplement the lectures in General Chemistry. The student attains a knowledge of elementary principles, becomes familiar with manipulating apparatus, and is prepared to commence analytical work.

The second year covers a systematic course in Qualitative Analysis in connection with the tests of the United States Pharmacopœia.

The third year is devoted to Volumetric Analysis by means of the standard solutions of the Pharmacopœia.



GENERAL CHEMISTRY AND PHYSICS

PROF GEORGE A. MENGE

LORING W. BEESON, Assistant

Inasmuch as the subjects of Analytical Chemistry and Pharmacy are fully provided for in other courses, these lectures will be devoted more closely to the fundamental principles and classification which must underlie a thorough and systematic knowledge of Chemistry. Owing to the intimate connection existing between Chemistry and several branches of Physics, a number of lectures illustrating the more important laws and principles of some of these branches will precede the regular course in Chemistry; and others on these subjects with which a prior acquaintance is less important will follow after the student has acquired some knowledge of chemical changes. This division of the lectures in Physics is furthermore necessary, in order that the student may not be too much handicapped in his preliminary laboratory work by a lack of familiarity with the simple facts and theory necessary to success therein.

During the first and half of the second year Physics and the non-metallic elements will receive consideration, followed during the remainder of the second and a portion of the third year by the Metals. The greater portion of the third year will be devoted to the exposition of the more important facts, principles, and theories of Organic Chemistry, including descriptions of the more important compounds; as well as of the larger classes and groups, both of the fatty and the aromatic series, which latter are gaining an ever wider use and importance in Pharmacy.

The value of an acquaintance by the pharmacist with the simplest elements of Physiological Chemistry is so great that two or three lectures devoted to their elucidation are given at the close of the third year.

Quizzes will be heard frequently throughout the three years, in order to test the students and to enliven their interest in the various subjects.

MICROSCOPY

PROF. BURTON J. HOWARD

The course of Microscopy is intended to give instruction in the use of the compound microscope as an aid in the study and identification of drugs. Attendance upon this course is required from Junior and Senior students.

The work will consist of both lectures and laboratory work, the exact details connected with each section of the plan of instruction will be clearly set forth as the course progresses.

The first part of the instruction will include the proper manipulation of the microscope and such features concerning microscopic technique as will be most necessary to students of Pharmacy.

After these studies the work will consist of the examination of plant tissue as illustrated in various vegetable substances most familiar to pharmacists. Special attention will be given to the structural characteristics by which one drug can be distinguished from another as well as the detection and identification of the most common adulterants used.

The importance of a knowledge of starches as a means of identification will involve a study of the most common ones, and will constitute an important feature in the first part of the course. The various parts of the plant organism will then be studied microscopically, using drugs as the basis of material. This will require in some cases the use of fresh material while in others dried samples will be used. Drugs commonly used in the powdered condition will first be examined in the entire state by means of sections, after which the powdered form will be studied.



MERCANTILE PHARMACY

PROF. HENRY B. FLOYD

It is generally conceded that many graduates in pharmacy are deficient in the qualifications necessary to conduct successfully the commercial side of their chosen profession. To overcome this condition, as much as is possible, a course of instruction will be given the Senior Class with a view of properly educating its students in business methods and their specialized application to the drug business.

In addition to the lectures herein briefly outlined, the course will cover the use and application of the fundamental principles of modern practical book-keeping, especially adapted to the needs of the pharmacist. Opening and closing books, calculating present worth, rapid and accurate posting, the making of trial balances and

statements, bills and other commercial paper, business letter writing and other important subjects will be taught by actual class room practice.

The lectures and practical work will include:

Book-keeping; single and double entry, the practical use, advantages and disadvantages of each system; the forms and books of modern book-keeping, the use and application of each; card, loose-leaf and book systems defined and explained, advantages and practical use of each form; the value and absolute necessity for a thorough system of book-keeping.

Banks and banking; descriptions of the several kinds of banks, individual accounts, business accounts, deposits, checks, collections, etc.

Partnership and incorporations, requirement, advantages and disadvantages of both.

Commercial paper; bills, notes, drafts, checks, etc., defined, their proper forms and uses explained, endorsements, etc.

Leases; responsibilities of landlord and tenant, the law of fixtures.

Contracts; definition, essentials of a valid contract, responsibilities of parties, redress, etc.

Methods of obtaining and establishing credit, allowance of retail credit, collections, etc.

Store system; ordering and marking goods; pricing; indexing; stock-taking; the necessity of inventory; the up-keep of stock; handling and filling orders; detection of leaks; etc.

Fire and life insurance; transportation; common and private carrier; responsibility of consignor and consignee; damages, etc.

Advertising; newspaper, window display, etc.

Value and use of mercantile agencies.

Business letter writing, forms, essentials, subject matter, etc.

Employee and employer, the responsibility and conduct of each.

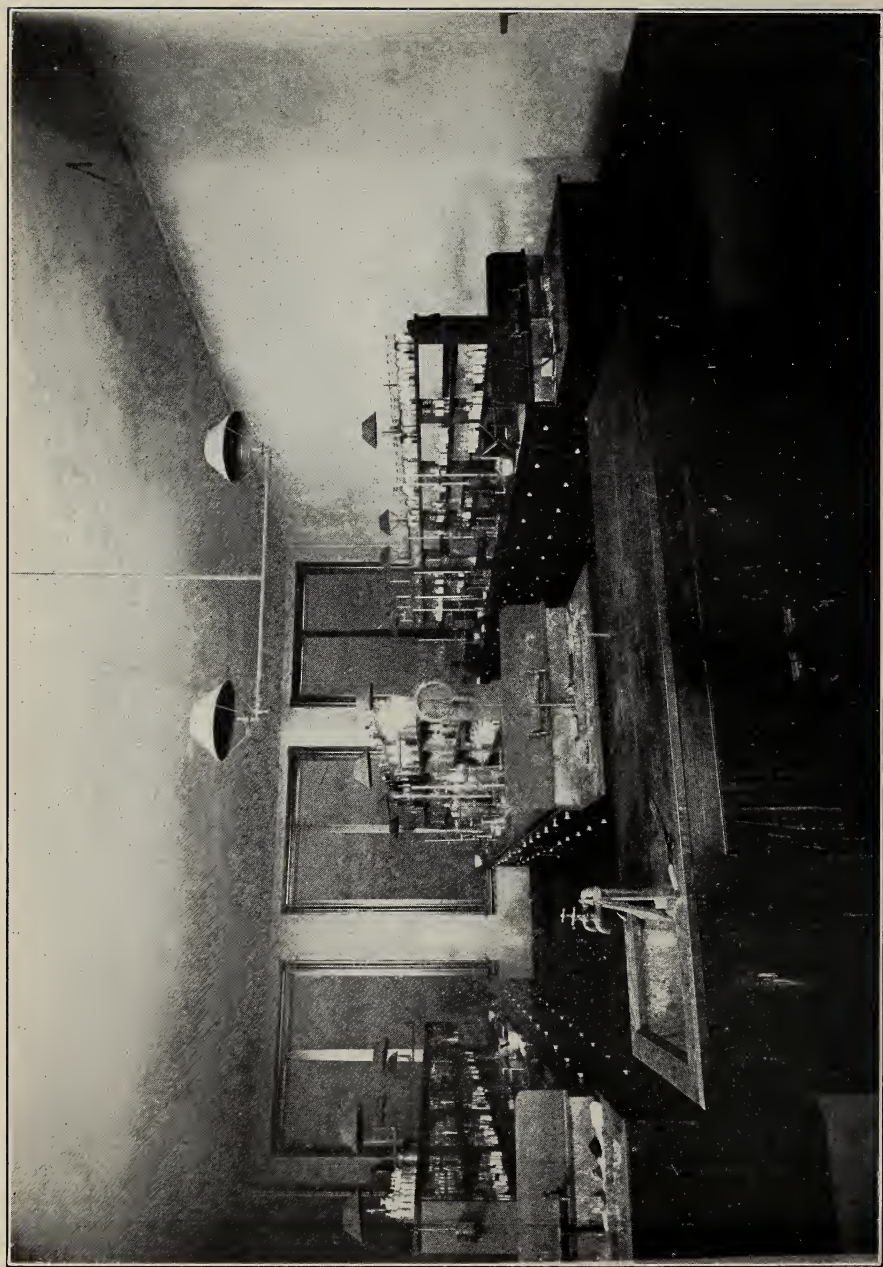


PHARMACAL JURISPRUDENCE PROF. ALEXANDER MUNCASTER

A course of twelve lectures to the Senior Class upon the rights and responsibilities of Pharmacists and the general and special laws bearing upon the practice of Pharmacy.

SYNOPSIS.—Historical sketch of the law in general; Constitution of the United States and State Constitutions; Federal law, with especial reference to the Food and Drugs Act of June 30, 1906, known as the "Pure Food Law," and special acts relating to Pharmacy in the District of Columbia; right to practice Pharmacy and the law in regard to prescribing; contracts in general and implied contracts of pharmacists; liability of retail, wholesale and manufacturing pharmacists for negligence; contributory negligence; statutes of frauds and limitations; corporations; partnerships; insurance; exemptions; bankruptcy; general business features, etc.

Students will be expected to take notes. The class will be quizzed during the course, and an examination held at the termination thereof.



A SECTION OF THE PHARMACEUTICAL LABORATORY

Syllabus of Lectures and Laboratory Instruction.

NOTE—Students should preserve this Syllabus as a guide to study

MATERIA MEDICA, BOTANY AND TOXICOLOGY

PROF. SAMUEL WAGGAMAN

FRESHMAN COURSE

1. Artificial and natural system of Botany. How plants grow and their essential parts.
2. Regular and irregular flowers; perfect and imperfect flowers. The theory of color, forms and character of the pollen. How flowers are fertilized.
3. The organs necessary for the perpetuation of plant life. The flower and its parts and its arrangement on the peduncle.
4. The various kinds of inflorescence: Raceme, corymb, umbel, thyrse, campanulate, cluster, head, compound, panicle, spike, spadix and catkin.
5. The general character of stems and the arrangement of the leaves on the petioles.
6. The difference between wild and cultivated flowers. The artificial means used by the florist to obtain variety of flowers and fruits.
7. The form and structure of ovules. The fruits which have a permanent calyx.
8. The classification and structure of fruits and their parts.
9. The seed and its parts, and how to distinguish the mono from the di-cotyledonous seed, and from what plants they are obtained.
10. The growth of the seed, and how it is distinguished from a fruit, and what fruits are improperly called seeds.
11. Tissues, and what part they play in the economy of the vegetable world; their simple and complex structure.
12. Cells and cellular tissue and the contents of cells; intercellular space.
13. Roots; their forms, structure and classification, and the best time to gather them for remedies.
14. Some unusual plants. The conditions under which plants may be entirely metamorphosed.
15. Buds and their peculiarities—flower buds, leaf buds, and mixed buds.
16. Leaves; their forms and structure.
17. The leaves; their surfaces, outlines and modifications.
18. What coloring matter leaves contain, and the best time to gather them for their medicinal virtues.
19. The parts and shapes of stems and their appendages. How to distinguish the endogens from the exogens.
20. Acrogens and their peculiarities. How fertilized.
21. Parasitic plants.
22. The mosses and liverworts.
23. Sea-weeds and their uses.
24. The fungi and a general resumé of the structure of plants as seen by the oxy-hydrogen light, and the importance of microscopic investigation of plant life.

PHARMACY

PROF. HENRY E. KALUSOWSKI

FRESHMAN COURSE.

1. Pharmacy defined. Description of Pharmacopœias and Dispensatories. Manner of reading formulæ. History of the United States Pharmacopœia. Method of revision.

2. Metrology. The principle and construction of various kinds of balances; their care, method of using, testing.

 Weighing. Standards of weight and measure; relation to each other.

 Measure of Capacity. How made and tested.

 Specific Gravity defined. Methods employed, illustrated by practical application to liquids and solids.

 Specific volume. Methods for determination illustrated and applied.

3. Heat. Sources from whence obtained; methods for applying; uses in Pharmacy; manner of regulating and modifying intensity by baths.

 Methods for Measurement of Heat. Manner of using thermometers. Relation of the various scales.

4. Evaporation, Principle of. Methods for conducting by heat, in vacuo, under pressure, spontaneously; effects of pressure and saturation of air; character of vessels affecting boiling point; rate and effects of evaporation and removal of products.

5. Distillation. Principle involved in the process. Construction of apparatus. Simple, fractional and destructive distillation. Manner of using distillatory apparatus. Recovery and treatment of distillates.

6. Fusion, Sublimation, Calcination, and Granulation. Principles explained and application to pharmacopœial uses described.

7. Desiccation and Comminution. Principles involved in construction of mills and cutters explained. Mortars and knives described. Garbling and powdering drugs. Sifting to required degree of fineness; considerations governing fineness of powders; effects of pulverization; pulverization of chemicals; commercial powders.

8. Solution. Theory of solution. Saturation. Solvents of the United States Pharmacopœia, classified and considered in groups.

9. Filtration. Explained and illustrated by means of various kinds of apparatus and media. Methods for rapid filtration by aid of heat or in the cold. Use of apparatus to produce vacuum; continuous filtration.

 Clarification and Decolorization. Methods for and substances used.

10. Precipitation. Theory of changes that take place and conditions required. Decantation; use of siphon; washing and drying precipitates.

11. Crystalization, Systems of. Measurement of crystals; conditions under which crystalization takes place; purification of crystals.

12. Dialysis and Diffusion. Consideration of principles; methods of application and uses.

13. Extraction, General purposes of and common results.

14. Maceration, Decoction and Infusion. Methods for effecting solution described. Official decoctions and infusions. Expression. Purpose of and methods described.

15. Percolation. Theory and principles; manner of preparing drug, menstruum, percolator and receiver; recovery of percolate; treatment of weak percolates; modified methods for percolation; repercolation described and illustrated.

16. Preparation of aqueous solutions of oils, viscid, mucilaginous, and saccharine substances; cold process for syrups described and illustrated. Official syrups, methods for preparing, description, character of and causes of deterioration.

17. Solutions, Mixtures, and Emulsions. Manner of making and character of compounds described. Glycerites and oleaginous solutions described; manner of preparation illustrated. Vinegars, resins and oleoresins; methods for preparation usually followed illustrated; character and cause of deposits in oleoresins.

18. Alcoholic, Hydro-Alcoholic, Vinous. Ethereal solutions described and methods for preparation demonstrated.

Extracts. Selection of drugs and menstrea; methods of concentration; treatment of reserved, concentrated and finished portions; condition and treatment of weak percolates by disillation.

Reviews at the close of the first half of the course and at the end of the last half of the course.

PHARMACEUTICAL LABORATORY WORK

FRESHMAN

The plan of instruction in these courses is arranged so as to give each student a thorough and practical knowledge of the various Pharmaceutical Processes, the application of which will be facilitated by an ample supply of material and apparatus to properly carry out the purpose of this course. That the work may be carried on according to a definite plan, it has been divided into sections, each of which is intended to be as nearly complete as the character of the process indicated will admit. *Students will be expected to complete the requirements of each section and submit the results of their work to the instructors for approval before they will be permitted to take up any work on the succeeding section.*

1. Students will receive instruction upon the proper use of the Pharmacopœia. By a series of written exercises upon subjects that will be named they will be required to write a description of a number of substances; the order of writing will include the official names, chemical symbols, synonyms, names of the parts or parts used; if a plant, name of natural order, description of substance named, solubility, melting point, tests for identification and names of official substances derived therefrom.

2. Testing balances, weighing and measuring; thermometers; rules for converting from one scale into that of another.

3. Specific Gravity. Methods for taking specific gravity of solids and liquids.

4. Application of heat to determine melting and boiling points, with arrangement of necessary apparatus to carry out a series of trials.

5. Preparation of series of aromatic waters by solution in hot and cold water filtering or otherwise clarifying the same.

6. Distillation. Fitting up distillatory apparatus and applying same for preparing distilled waters, recovering alcohol, and for the preparation of not less than two solutions of gases in water and two aromatic waters.

7. Solution. Applied to the preparation of decoctions and infusions and of syrups mucilages, vinegars, and glycerites.

8. Precipitation, exsiccation and granulation applied to a number of inorganic salts.

9. Percolation, applied to the preparation of a series of products, among which will be included aqueous extracts, hydroalcoholic extracts and tinctures, with methods for recovery and treatment of alcohol from weak percolates. The preparation of a series of syrups by the "cold process," or percolation, will follow.

10. Upon the conclusion of the foregoing, a laboratory examination will be held; this will involve the application of any of the above named processes to the preparation of some product that will be named on the date of the examination.

INTRODUCTORY LABORATORY PRACTICE.

PROF. HOWARD M. BRADBURY

FRESHMAN COURSE.

1. Phenomena of attraction: Gravitation, weight, center of gravity, cohesion, adhesion, capillarity.
2. Heat: Mechanical motion and heat convertible; temperature; conduction, convection, radiation. Expansion in solids, in fluids; coefficient of expansion; laws of expansion of gases; kinetic theory of gases.
3. Heat *continued*. Specific heat, calorimetry; change of state produced by heat, liquefaction, vaporization, distillation, sublimation. Pressure of vapors. Sources of heat. Dynamical theory of heat.
4. Review.
5. Examination in Physics.
6. Care of apparatus; rules to be observed in laboratory work; importance of recording results and method of recording same; evidences of chemical change and distinction from physical change.
7. Study of the properties of oxygen and methods of preparation. Oxygen produced by heating oxides, such as mercuric oxide and manganese dioxide.
8. Preparation of hydrogen by action of metals on water. Zinc and sulphuric acid. Electrolysis of water. Union of hydrogen with oxygen to form water.
9. Manufacture of soap. Hard and soft water.
10. Experiments with substances containing water of crystallization, and with deliquescent bodies. Hydrogen peroxide.
11. Oxygen continued. Preparation from salts containing oxygen, such as potassium chlorate. Study of the law of definite proportions as illustrated by these reactions, calculations by means of atomic and molecular weights.
12. Reduction of copper oxide by hydrogen. The synthesis of water.
13. Review.
14. Examination in Chemistry.
15. Chlorine. Preparation by the agency of an oxydizing body on a chloride, such as manganese dioxide and sodium chloride in presence of sulphuric acid; properties of chlorine. Preparation from sodium chloride by electrolysis of salt solution, illustrating the commercial production of chlorine by the use of the electric current. Preparation by oxidation of hydrochloric acid with potassium permanganate. Preparation of hydrochloric acid by means of sodium chloride and sulphuric acid.
16. Compounds of chlorine, oxygen and a metal. Action of chlorine on alkali hydrates. Chlorates. Hypochlorites made by electrolysis. Bleaching powder.
17. Acids, bases, salts. Neutralization. Definite quantities of acids and bases unite to form salts. Test papers. Salts formed by the action of acids on metals and oxides.
18. Review.
19. Air. Oxygen and nitrogen in air. Moisture shown by deliquescence. Carbonic acid in air. Chemical changes produced by air.
20. Ammonia. Destructive distillation of substances containing nitrogen. Salts of ammonium in presence of alkalies.
21. Compounds of nitrogen, oxygen and hydrogen. Nitric acid prepared from potassium nitrate and sulphuric acid. Oxidation by nitric acid.
22. Review.
23. Carbon. Occurrence. Amorphous carbon. Carbon as a reducing agent.
24. Compounds of carbon. Carbon dioxide, preparation and detection. Absorption of carbon dioxide by solution of alkali hydroxides. Carbon monoxide by decomposition of oxalic acid. Bunsen burner.

25. Alkali metals. Preparation of sodium carbonate by the LeBlanc method.
26. Alkali metals. Preparation of caustic soda by lime; electrolytic production of caustic soda and chlorine from common salt. Ammonium hydroxide from crude ammonia liquor.
27. Study of elements in groups. Periodic system. Chlorine group.
28. Study of elements in groups continued. Calcium group.
29. Review.
30. Review.

The result of all experiments are required to be recorded in note books, which are criticised each week.

GENERAL CHEMISTRY

PROF. GEORGE A. MENGE

FRESHMAN COURSE

1. Physical and chemical phenomena. Physical agents. Matter. Molecular theory. Properties of matter. Force. Molar and molecular forces. Different states of aggregation.
2. Dynamics of liquids. Compressibility; pressure and equilibrium; buoyant force; density, capillarity, diffusion, osmosis, dialysis siphon; pumps.
3. Pneumatics: Physical properties of gases, the atmosphere, barometer, compressibility and expansibility of gases, air pump. Gas laws and their application. Diffusion, absorption.
4. Review.
5. Examination in Physics.
6. Chemical change; relations between chemical and physical change. Definition of chemistry. Composition of matter; elements, mixtures, compounds. Chemical action, affinity. Reaction. Analysis, synthesis. The chemical elements and their symbols.
7. Oxygen: Occurrence, preparation. Symbolic expression of reactions, chemical equations. Properties of oxygen. Oxidation, combustion, respiration. Heat of combustion and of decomposition. Chemical energy and chemical work. Oxides.
8. Hydrogen: Occurrence, preparation. Properties of hydrogen, diffusive and penetrative power, diffusion of gases in general, reducing power of hydrogen; combination with oxygen, oxhydrogen blow-pipe.
9. Water: Formation and proof of its composition. Measurement of gases. Properties of water. Consequences of its abnormal expansion and contraction. Expansive force of freezing water. Ice.
10. Water *continued*: Solvent power, mineral waters, sea water. Solution. Water of crystallization, efflorescence, deliquescence, hygroscopicity. Hydrogen dioxide.
11. Conservation of mass and energy. Laws of definite and multiple proportions. Atomic theory. Atomic and molecular weights. Formulas, their meaning and use.
12. Study of reactions employed in the preparation of oxygen and hydrogen and in the study of water. Classes of chemical reactions.
13. Review.
14. Examination in Chemistry.
15. The halogens. Chlorine: Occurrence, preparation, properties. Hydrochloric acid.
16. Oxides and oxyacids of chlorine; hypochlorites, chlorates.
17. Acids, bases, neutralization, salts. Chemical nomenclature of compounds. Electrolytes and non-electrolytes.
18. Review.
19. Nitrogen: Occurrence, preparation, properties. The atmosphere, composition, analysis, function.

20. Nitrogen *continued*: Ammonia, its formation and manufacture, properties, uses, composition. Relations between the volumes of combining gases. Relations between the specific gravities of gases and their combining weights. Ammonium salts. Compound radicals.

21. Nitrogen *continued*: Oxides of oxyacids of nitrogen. Nitric acid; its manufacture, properties, uses; fuming nitric acid; aqua regia. Nitrous oxide, nitric oxide, nitrogen trioxide and nitrous acid, nitrogen tetroxide. Chlorides and iodides of nitrogen.

22. Review.

23. Carbon: Occurrence. Allotropy. Diamond, graphite, amorphous carbon. Chemical properties of carbon, reducing power. Compounds with hydrogen. Organic chemistry; petroleum, methane, ethylene, acetylene.

24. Carbon *continued*: Carbon monoxide; carbon dioxide; its occurrence, preparation, properties. Respiration. Carbonic acid and carbonates.

25. Illuminating gas, combustion, flame. Bunsen burner, blow-pipe, miner's lamp. Cyanogen, hydrocyanic acid. Carbon in metallurgy.

26. Review.

27. Avogadro's hypothesis. Relation between specific gravity and molecular weights of gases. Elemental molecules, nascent state. Molecular equations. Determination of formulas. Valence; replacing power of elements.

28. Methods of determining molecular weights and atomic weights. The Mendeleeff classification of the elements. Principal elements in each group.

29. Review.

30. Review.

MATERIA MEDICA, BOTANY AND TOXICOLOGY.

PROF. SAMUEL WAGGAMVN

JUNIOR COURSE.

1. Althea, pyrethrum, apocynum, asclepias, and belladonna roots. Order and habitat.

2. Berberis, chondodendron tomentosum, euphorbia, sumbul, frazera, gentiana, glycyrrhiza. Order and habitat.

3. Ipecacuanha, jateorrhiza, krameria, and lappa roots; their order and habitat.

4. Lovage, apium, phytoacca, senega, rhenm, smilax, saponaria, and sassafras roots. Order and habitat.

5. Scammonium, selinum, stillingia, symphytum, arnica, asarum, felix-mas, gelsemium roots. Order and habitat.

6. Geranium, cimicifuga, helleborus, hydrastis, iris, veratrum, leptandra, podophyllum, and rubia tinctora roots. Order and habitat.

7. Sanguinaria, serpentaria, spigelia, tritium, tormentilla, valeriana, zingiber. Order and habitat.

8. Aconite, allium, arum, colchicum, corydalis, jalapa, scilla. Order and habitat

9. Dulcamara, juniperus, sabina, scoparius, thuja, guaiacum, hematoxylon, quassia, santalum. Order and habitat.

10. Angustura, azedarach, cinchona, cinnamon, and wahoo barks. Order and habitat.

11. Rhamnus, gossypium, juglans, mezereon, nectandra, prinos, prunus, punica-granatum, and quillaia. Order and habitat.

12. Quereus, rubus, salix, ulmus, viburnum, and xanthoxylum. Order and habitat.

13. Aconite, aurantium, buchu, cannabis, chimaphila, digitalis, erythroxyllum, hyoscyamus, kalmia, matico, and lycopus leaves. Order and habitat.

14. *Myrcia*, *pilocarpus*, *rosemary*, *stramonium*, *senna*, *tabacum*, *uva ursi*, *thea*. Order and habitat.

15. Herbs: *Absinthium*, *achillea*, *epigea*, *erigeron*, *eupatorium*, *gaultheria*, *grindelia*, *hedeoma*, *hyssopus*, *lobelia*, *marrubium*, and *melissa*. Order and habitat.

16. Herbs *continued*. *Mentha*, *nepeta cataria*, *origanum*, *ranunculus*, *sabatia*, *scutellaria*, *solidago*, *tanacetum* and *thymus*. Order and habitat.

17. Flowers and petals, *aurantium*, *chamomile*, *caryophyllus*, *rosa*, *santonica*, *kooso*, and their essential properties and adulterants.

18. Fruits, *anethum*, *anisum*, *apium*, *capsicum*, *cardamomum*, *carum*, *carota*, *cassia fistula*, *citrus*, *chenopodium*, *cocculus*, *colocynthis*, and *coriandrum*. Order and habitat, active principles and doses.

19. Fruits *continued*. *Cubeba*, *cuminum*, *fœniculum*, *macis*, *myristica*, *papaver*, *pimento*, *piper*, *rhus glabra*, and *vanilla*. Order and habitat.

20. Seeds, *avena*, *colchicum*, *almond*, *cydonium*, *ignatia*, *nux vomica*, *tonca*, *linum*, and *sinapis*.

21. Excrescences, piths, hair glands, starches, and sugars:

22. Balsams, oleoresins, gum and gum resins.

23. Fixed and volatile oils.

24. Volatile oils, waxes and animal drugs.

PHARMACY

PROF. HENRY E. KALUSOWSKI

JUNIOR COURSE

1. Chlorine, Bromine, Iodine, Phosphorous and Sulphur. History, sources, process of production, methods of preparation of officinal compounds, tests for identification.

2. Carbon, Boron and Silicon. Compounds used in Pharmacy and the methods of production.

3. Acids. Preparation, tests, impurities, purification, and pharmacopœial uses; use of hydrometers and specific gravity tables for determining strength thereof.

4. Potassium, Sodium, Lithium and Ammonium. Sources of supply, methods of producing and purifying, pharmacopœial compounds, tests for identity, and composition.

5. Magnesium, Calcium, and Barium. Sources of supply, production of compounds, methods for purification, tests and uses in Pharmacy.

6. Zinc, Aluminum, Cerium and Cadmium. Sources of supply, methods of obtaining the elements, formation of compounds, impurities, tests for identity and composition.

7. Manganese, Iron and Chromium. Methods for obtaining pharmacopœial compounds from these elements, characteristics of, processes of purification, tests for identity and composition; scale salts of iron, their character and preparation.

8. Lead, Copper, Silver and Mercury. Sources of supply, manner of production, characteristics of; pharmacopœial compounds obtained from these elements, tests for purity and identification.

9. Antimony and Arsenic. Sources, production of salts, methods for testing. Toxicology; antidotes and methods of administration.

10. Bismuth. Preparation of pharmacopœial compounds; impurities, separation of; tests.

11. Gold. Notice of officinal compounds; methods for testing.

12. Fixed Oils and Fats; their composition, methods for production and purification, keeping and dispensing. Tests for adulteration, saponification and iodine absorption values demonstrated.

13. Volatile Oils, Sources. Composition of the more important kinds in the light of recent investigations. Tests for identity and assay methods.

14. Alkaloids and Glucosides, Sources of. Methods for separation, character, test^s for identification, remarks on dispensing and antidotes, followed by drugs containing neutral, cathartic and astringent principles.

15. Animal Products. Fats, oils, lactic acid, anti-diphtheric serum. Powdered glandular substances. Pepsin. Pancreatine. Preparation and methods of testing.

Reviews at the close of the first half of the course and at the end of the last half of the course.

PHARMACEUTICAL LABORATORY WORK.

JUNIOR.

1. Preparation and standardization of the following solutions:

Half normal hydrochloric acid solution.	
Normal potassium hydrate solution.	Decinormal silver nitrate solution.
Half normal alcoholic potass hydroxide	Tenth normal potass permanganate solution.
Decinormal oxalic acid V. S.	Mercuric chloride T Sol.
Decinormal iodine V. S.	Decinormal sodium hyposulphite solution.
Normal sulphuric acid.	

Quantitative and qualitative tests of the following inorganic compounds will be made:

2. Acids. Hydrochloric, Nitric, Phosphoric, Sulphuric, Acetic, Tartaric and Citric.
3. Chlorinated lime, Iodine.
4. Potassium and Sodium hydrates, Potassium bitartrate and iodide, Potassium and sodium tartrate.
5. Sodium, bicarbonate, borate, bromide and phosphate. Ammonia water.
6. Iron reduced. Sol. Chloride iron, Citrate iron and quinine.
7. Solution Subacetate lead, Silver nitrate.
8. Antimony and potassium tartrate, Arsenous acid.
9. Determination of the purity of a series of fixed oils and fats by chemical tests, finding of saponification numbers and iodine absorption values.
10. Tests for purity and identification of acetanilid, phenacetine, sulfonal and such other "synthetics" as may be given.
11. Quantitative assays of opium, nux vomica, Belladonna and preparations of alkaloid containing drugs.
12. Assays of pepsin and pancreatin.
13. Upon the conclusion of the foregoing a laboratory examination will be held. It will consist of problems involved in any of the above numbered sections; the problems will be announced on the date of the examination.

ANALYTICAL CHEMISTRY.

QUALITATIVE ANALYSIS.

PROF. HOWARD M. BRADBURY.

JUNIOR COURSE.

1. Qualitative analysis; description of operations and apparatus used in analysis; terms defined.
2. Classification of the elements for qualitative analysis; classification of the elements for other purposes.
3. Bases divided into groups by means of reagents. Alkali group: Potassium, sodium, ammonium, lithium.
4. Symbols, formulæ, and equations used in recording the results of analysis. Preliminary reactions of the calcium group: Barium, calcium, strontium and magnesium.
5. Calcium group continued. Equations explanatory of the chief reactions; study of the group by the scheme of analysis.
6. Iron and zinc groups. Preliminary reactions with reagents, such as alkali hydrates, carbonates, sulphides, phosphates, &c., with the metals of this group.
7. Iron and zinc group continued. Writing equations under this group; study of the group by means of the scheme of analysis.

8. Iron and zinc group continued. Analysis and identification of the members of the group in presence of phosphates and interfering substances.

9. Copper group, division B. Preliminary reactions with test substances; writing equations.

10. Copper group, division B, continued. Study of group by the scheme of analysis.

11. Tin group (arsenic, antimony and tin). Preliminary reactions with test substances; study of the methods of separating arsenic and antimony.

12. Tin group continued. Separation from the copper group; analysis and identification by the scheme; equations.

13. Silver group. Tests and analysis by the scheme; equations.

14. Separation of the groups of metals and the identification of members in each group.

15, 16, 17. Continued drill on the use of the scheme of analysis for the detection of unknown elements.

18. Acid-forming elements. Acids—monobasic, dibasic, tribasic; theory of their structure; tests for sulphides, carbonates, nitrates, sulphates, sulphites.

19. Analysis of salts of unknown metals combined with hydrosulphuric, carbonic, nitric, sulphuric and sulphurous acids.

20. Acids continued. Tests for halogen acids—hydrochloric, hydrobromic, hydriodic—and their separation.

21. Analysis of salts of unknown metals combined with the halogen acids.

22. Halogen acids continued. Tests for hydrocyanic, hydroferro- and hydroferri-cyanic acids and their separation from each other.

23. Analysis of salts containing metals in combination with preceding acids.

24. Hydrochlorous, chloric, boracic and hydrofluoric acids.

25. Acids of phosphorus, hypophosphorous, phosphoric, pyro- and metaphosphoric acids.

26. Final review of the schemes of analysis.

27. Practical laboratory examination.

GENERAL CHEMISTRY.

PROF. GEORGE A. MENGE.

JUNIOR COURSE.

1. Light; its nature, sources, velocity; reflection; mirrors; refraction; total reflection; prisms; spectra; etc.

2. Light continued: Lenses; images; chromatic and spherical aberration; compound microscope; undulatory theory; color phenomena.

3. Magnetism: its nature, production, peculiarities and uses. Frictional and dynamical electricity; induction, potential, electro-chemical decomposition, electro-chemical series, cell, battery, etc.

4. Dynamical electricity continued: Electrolysis and the ion theory; secondary batteries; electrical measurements; dynamos; physiological effects of electricity application in medicine.

5. Review.

6. Review the Mendeleeff classification. The nonmetals.

Group VII. Bromine, hydrobromic acid, oxyacids of bromine. Iodine, hydriodic acid. Fluorine, hydrofluoric acid. Comparison of the members of the group.

7. Sulphur Group: Sulphur; occurrence, allotropic forms, properties, uses. Crystallography, polymorphism.

8. Sulphur continued: Hydrogen sulphide: Oxides and acid of sulphur. Sulphurous oxide and acid. Sulphuric acid; its manufacture, purification, properties.

9. Sulphur continued: Monobasic, dibasic and polybasic acids. Acid, neutral and normal salts. Other acids of sulphur. Carbon bisulphide. Selenium, tellurium. General remarks on the sulphur group.

10. Review.
11. Group V: Phosphorus; occurrence, allotropic forms, properties. Hydrides. Hypophosphorous acid and hypophosphites. Phosphorous oxide and acid. Phosphoric oxide and acids.
12. Arsenic; occurrence, properties. Hydrides, oxides, oxyacids, sulphides, etc. Detection of arsenic in cases of poisoning.
13. Group IV: Review Carbon. Silicon: Occurrence, etc. Silica and silicic acid; silicates, their role in nature.
- Group III: Boron; occurrence, etc; boric oxide and acid; borax, properties and uses.
14. Review.
15. Metals: General characteristics, common occurrence, physical properties, specific heat; atomic heat, etc. Isomorphism. Alloys.
16. Group I. The alkali metals: Potassium, occurrence, preparation, properties; Its salts, their manufacture and uses.
17. Sodium: Preparation and properties; its salts, their manufacture and uses.
18. The radical ammonium; its salts, etc. Lithium, cesium, rubidium. Review of the group.
19. Review.
20. Group II. The Alkaline earth metals. Calcium; properties of the metal and its chief salts. Lime, plaster of Paris, gypsum, phosphate fertilizers.
21. Strontium and barium; their chief salts. Review of the group. Spectrum analysis.
22. The Magnesium group. Magnesium; zinc; preparation and properties of the metals and their common salts.
23. Cadmium; Mercury; preparation and properties of the metals and their common salts.
24. Review.
25. Copper, Silver and Gold: Their metallurgy, properties, important salts, uses, etc. The platinum metals.
26. The Aluminum group: Aluminum, occurrence manufacture, properties, uses, salts, etc. The alums, clays, ultramarine. Rare earth metals.
27. The iron group. Iron; its ores and metallurgy; cast and wrought iron, steel. Salts of iron; their properties and uses. Nickel and cobalt.
28. Review.
29. Review.

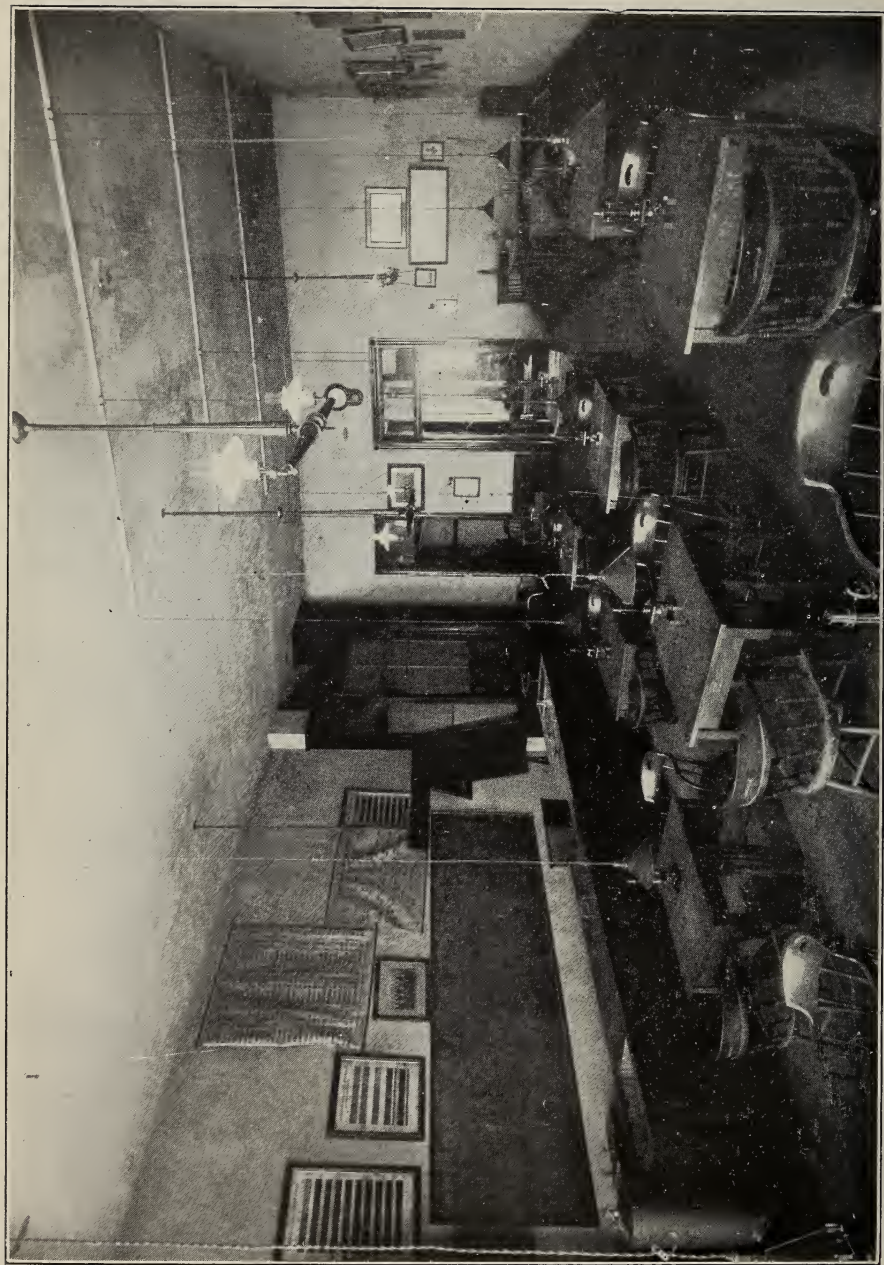
MICROSCOPY.

PROF. BURTON J. HOWARD.

JUNIOR COURSE.

[NOTE.—The following syllabus will give some idea as to the character and order of studies for the Junior Course, though it may be found necessary to change the order as well as some of the illustrative matter:]

1. Optics of microscopes: Refraction of light, simple lenses, chromatic aberration, spherical aberration.
2. Forms of simple magnifiers. The compound microscope. Types of objectives.
3. Numerical aperture, mechanism of compound microscopes, micrometry of simple lenses, of compound microscopes.
4. Laboratory tests for focus, determination of magnification in magnifiers, in compound microscopes.
5. Micrometry work.
6. Study of starches: Wheat, potato, arrowroot, corn, rice, cassava, rye.



A SECTION OF MICROSCOPICAL LABORATORY

8. Woods: Elements constituting woods, in angiosperms, in gymnosperms.
9. Woods: Quassia, yellow sandal wood.
10. Microchemical tests: For cellulose, for lignified tissue.
11. Types of calcium oxalate crystals.
12. Stems: Lobelia stem, dulcamara stem.
13. Leaves: General histology of.
14. Leaves: Senna, buchu, tea, hyoseyamus, mullein.
15. Barks: General histology of.
16. Barks: Oak, cascara, cinnamon.

SENIOR COURSE.

1. Seeds: Structure, embryo, coverings, reserve materials.
2. Seeds: Mustard seed, linseed, coffee, nux vomica.
3. Fruits: Cardamon fruit, capsicum, pepper, wheat.
4. Rhizomes: General structures.
5. Rhizomes: Arnica, ginger.
6. Roots: Dandelion root, chicory, ipecacuanha.
7. Some common adulterants: Cocoanut shells, olive pits, cracker crumbs, turmeric.
8. Practical work: Consisting of practical work in the identification of mixtures of substances which have been studied during the Junior and Senior courses.

MATERIA MEDICA, BOTANY AND TOXICOLOGY.

PROF. SAMUEL WAGGAMAN.

SENIOR COURSE.

1. Roots; their constituents, active principles, adulterations, official preparations, therapeutic uses, doses and antidotes.
2. Rhizomes; their constituents, active principles, adulterations, official preparations, therapeutic uses, doses and antidotes.
3. Tubers and bulbs; their constituents, active principles, adulterations, official preparations, therapeutic uses, doses and antidotes.
4. Twigs and woods; their constituents, active principles, adulterations, official preparations, therapeutic uses, doses and antidotes.
5. Barks; their constituents, active principles, adulterations, official preparations, therapeutic uses, doses and antidotes.
6. Leaves and leaflets; their constituents, active principles, adulterations, official preparations, therapeutic uses, doses and antidotes.
7. Herbs; their constituents, active principles, adulterations, official preparations, therapeutic uses, doses and antidotes.
8. Flowers and petals; their constituents, active principles, adulterations, official preparations, therapeutic uses, doses and antidotes.
9. Fruits; their constituents, active principles, adulterations, official preparations, therapeutic uses, doses and antidotes.
10. Seeds; their constituents, active principles, adulterations, official preparations; therapeutic uses, doses and antidotes.
11. Excrecences, piths, hairs, glands and starches, official preparations, therapeutic uses, etc.
12. Juices, extracts, and sugars; their constituents, official preparations, therapeutic uses, etc.

16. Balsams and oleoresins; their constituents, active principles, adulterations therapeutic uses, doses.
17. Gums; their constituents, active principles, adulterations, therapeutic uses.
18. Gum resins and resins; their constituents, active principles, adulterations, therapeutic uses.
19. Fixed oils and waxes; their constituents, active principles, adulterations, therapeutic uses.
20. Volatile oils; their constituents, active principles, therapeutic uses, doses.
21. Animal drugs, as compared with vegetable drugs; therapeutic value.
22. The most recent remedies and their therapeutic value and doses. Mineral and corrosive poisons and their antidotes.

PHARMACY.

PROF. HENRY E. KALUSOWSKI.

SENIOR COURSE.

1. Cellulose Group. Products obtained directly and by destructive distillation. Coal-tar products under this head will be specially treated. A number of preparations classed under the general term of "synthetics," tests for identification and their purity will be considered.
2. Amylaceous, Mucilaginous and Saccharine substances as such, and their more important products, will be taken up and noted according to their value.
3. Derivatives of Sugars through action of ferments. Alcohols and Ethers. Methods by which they are obtained; purification tests for identification and purity.
4. Aldehyde; its derivatives, preparations and tests.
5. Products obtained by the action of ferments upon Acid and Saccharine Fruits.
6. Soaps. Methods and theory of saponification, purification of by-products.
7. Products obtained from Resins, Balsams and Gum Resins. Methods of separation and production in a state of pharmaceutical purity; tests.
8. Methods for producing Resenoids with remarks on their relations to the medicinal properties of the drug from whence they are derived and the system of their nomenclature.
9. Mixtures. Methods of preparing official and unofficial mixtures, emulsions of oils, gum resins; resins, liniments, and liquids that are insoluble in aqueous vehicles; explained and demonstrated.
10. Pills. Methods of making masses; hardness and coherence of the same; how regulated by excipients; choice of excipients and functions of absorbing powders; treatment of oils and other liquids in making masses; division and shaping mass into pills; finishing and dusting pills; choice of dusting powders.
11. Pill Coating, Sugar, Gelatine, Chocolate and Foils. Methods for and composition of coating materials.
12. Tablet Triturates. Composition; use of excipients; manner of division and methods for forming.
13. Capsules, Composition of. Manner of capsuling pills and powders; capsuling and dispensing liquids; solubility of capsules; conditions affecting the same; care in dispensing.
Wafers, Composition of. Methods for filling and dispensing.
14. Compressed Tablets. Methods followed in preparing the mass and compressing.

15. Troches, Masses and Confections. Medicaments adapted to that form of administration.

16. Powders. Methods for securing uniformity of diffusion in simple and compound powders; dividing, folding and dispensing.

Treatment of Powders. Composed in part or whole of volatile, efflorescent, or deliquescent substances.

17. Suppositories. Hot and cold processes; use of molds; preparation of the mass; incorporating the medicaments; rolling and dividing, shaping; use of molds explained and demonstrated.

18. Plasters. Methods for preparing and spreading. Remarks on dispensing.

19. Ointments. Cerates and other bases for applying medicaments by contact or inunction; methods for preparing; use of heat; necessity for straining and purification from foreign matter; manner of incorporating solution of salts, waters, extracts and powdered substances; preservation of products and methods of dispensing. Oleates. Methods for making by direct combination; solution or decomposition; remarks on their character and qualities.

20. Milk, composition of; seasonal variation. Methods for testing for fats, albuminous substances, sugar, water.

21. Urinary Analysis. General remarks on normal and pathological constituents of urine; variations from normal interpreted; description and application of chemical tests; application and use of the microscope for examining urinary sediments; collection of sediments and methods for mounting described and demonstrated.

Review at the close of the first half of the course and at the end of the last half of the course.

PHARMACEUTICAL LABORATORY WORK.

SENIOR.

1. Percolation, Various methods of. Applied to making a series of fluid, solid extracts and powered extracts.

2. Recovery of alcohol from exhausted drugs and weak percolates; determination of the percentage of alcohol in the recovered liquid; instruction for converting it into diluted alcohol of other required strength.

3. Solution. Applied to the preparation of four or more solutions of iron salts with assay for percentage composition.

4. Pills, Powders, Capsules, Wafers, Tablet Triturates. Students in this section will be expected to make up a series of these preparations, and will be instructed in the manner of making and methods of dispensing. Pill coating will receive due attention. Extemporaneous mixtures will close the work in this section.

5. Granular Effervescent Salts. Methods for preparing and granulating. Compressed Tablets. Methods for preparing the mixture and compressing.

6. Continuing through this course instruction in prescription work will be given. It will include methods for compounding, and the attention of students will be directed to the details involved in dispensing, problems in deciphering illegible prescriptions, and the treatment of incompatibles, whether chemical, physical, or therapeutical, will receive the attention that their importance justifies. Students will be furnished with copies of prescriptions for compounding. The prescriptions will be selected with the view of presenting difficulties and incompatibilities ordinarily met with in the practice of Pharmacy.

7. Preparation of Scale Salts. This will begin by taking definite quantities of solution of tersulphate of iron, and following each step to the production of the finished scale.

8. Suppositories, Oleates, Plasters, Ointments and Cerates. Instruction in preparation and methods for dispensing.

9. Upon the conclusion of the work indicated in each section a laboratory examination will be held. It will involve the production of some compound or compounds included in any of the sections; the nature of the product or products required will be announced on the date of examination.

ANALYTICAL CHEMISTRY.

QUANTITATIVE ANALYSIS.

PRO. HOWARD M. BRADBUBY.

SENIOR COURSE.

1. Quantitative analysis; gravimetric; volumetric; standard and normal solutions; indicators, theory and use of.
2. Apparatus used in volumetric analysis; care, use, methods of calibration and corrections applied.
3. Theory of solution. Principles of chemical combination as applied to problems in volumetric analysis.
4. Alkalimetry. Estimation of solution of potassium hydroxide and ammonia water by means of normal sulphuric acid V. S.
5. Estimation of samples of sodium carbonate by hot process. Direct titration.
6. Estimation of samples of sodium carbonate by residual titration employing empirical alkali solutions.
7. Estimation of ammonium carbonate by means of normal sulphuric acid V. S.
8. Standardization of hydrochloric acid by sodium carbonate.
9. Estimation of potassium and sodium tartrate by residual titration demonstrating the use of standard solutions not strictly normal or systematic.
10. Preparation of normal potassium hydroxide V. S., including standardization against potassium bitartrate as set forth in the pharmacopoeia.
11. Acidimetry. Estimation of diluted hydrochloric acid and sulphuric acid by means of normal potassium hydroxide V. S.
12. Analysis by precipitation. Estimation of potassium iodide and potassium cyanide by tenth-normal silver nitrate V. S.
13. Analysis by precipitation continued. Estimation of ammonium bromide in presence of chlorides by tenth-normal silver nitrate V. S.
14. Standardization of hydrochloric acid with silver nitrate.
15. Analysis by oxidation. Preparation of tenth-normal potassium permanganate V. S. and its standardization with metallic iron.
16. Preparation of tenth-normal potassium permanganate V. S. by mixing two solutions of different strengths in proportion based upon the quantity of each required to react with a definite quantity of tenth-normal oxalic acid V. S.
17. Estimation of ferrous sulphate by tenth-normal potassium permanganate V. S.
18. Estimation of solution of hydrogen dioxide for weight of absolute hydrogen dioxide and volume of oxygen.
19. Determination of ferrous and ferric iron in a solution of salts of each by means of tenth-normal potassium dichromate V. S.
20. Indirect oxidation or iodometric estimations. Preparation of tenth-normal iodine V. S. and its standardization with arsenic trioxide.
21. Preparation of tenth-normal sodium thiosulphate V. S. and its standardization against tenth-normal iodine V. S. Estimation of chlorine water.
22. Estimation of Labarraque's solution by means of tenth-normal sodium thiosulphate V. S.
23. Estimation of solution of ferric chloride by means of tenth-normal sodium thiosulphate V. S.
24. Estimation of sugar by Fehling's solution. Determination of glucose in syrups.
25. Estimation of solution of formaldehyde by means of fifth-normal potassium permanganate V. S., using residual titration method with empirical solution of peroxide of hydrogen (Free Zeit 40-587).
26. Practical laboratory examination.



A SECTION OF CHEMICAL LABORATORY

GENERAL CHEMISTRY.

PRO. GEO. A. MENGE.

SENIOR COURSE.

1. Manganese and chromium. Occurrence, ores, salts. Molybdenum, tungsten, uranium.
2. Tin and lead, their ores, metallurgy and salts. Allied elements.
3. Antimony and bismuth. The Periodic system.
4. Organic chemistry. Definition. General properties of organic compounds, their sources, purification and analysis. Determination of boiling and melting points. Rational and structural formulas, isomerism, classification.
5. The hydrocarbons—methane and ethane. Homologous series. Halogen derivatives of methane and ethane: Oxygen derivatives: Alcohols.
6. Methyl and ethyl alcohol. Fermentation. Ethers. Ethyl ether, mixed ethers.
7. Aldehydes of methane and ethane. Chloral. Acids: Formic and acetic acids; acid anhydrides. Relations between acids, alcohols, aldehydes and hydrocarbons.
8. Esters or ethereal salts. Ketones. Sulphur derivatives of methane and ethane.
9. Nitrogen derivatives of methane and ethane: Cyanogen, hydrocyanic acid, ferro- and ferri-cyanides, cyanides and isocyanides, cyanates and isocyanates.
10. Substituted ammonias; amines or amino compounds. Nitro compounds. Metallic derivatives.
11. The marsh gas series of hydrocarbons or paraffins. Petroleum, its refining. Isomerism among the paraffins.
12. Oxygen derivatives of the higher members of the paraffin series; characteristics of normal, secondary, and tertiary alcohols. Higher aldehydes and fatty acids; soaps.
13. Polyacid alcohols and polybasic acids; oxalic acid, etc.; glycerine, fats, saponification.
14. Mixed derivatives of the paraffins. Hydroxyacids, glycolic, lactic, malic, tartaric, citric, saccharic, mucic acids, etc.
15. Carbohydrates: Glucoses, saccharoses; celluloses, gums. Polarization.
16. Mixed compounds containing nitrogen. Amido acids, acid amides. Unsaturated carbon compounds; ethylene, acetylene.
17. Aromatic compounds: The benzene series of hydrocarbons and derivatives. Theory of the benzene ring. Isomerism among benzene derivatives. Benzene, toluene.
18. Nitrobenzene, aniline, toluidine, diazo-compounds. Phenols and cresols, picric acid, creasote, thymol, etc. Benzoic aldehyde, etc. Benzoic and salicylic acids, etc. Naphtalene, anthracene.
19. Pyridine bases, terpenes, camphors, glucosides, vegetable and animal alkaloids, proteids.
20. Physiological chemistry: Chemical changes in plants and animals.
21. Physiological chemistry *continued*. Animal fluids and tissues.

MEMORIAL SCHOLARSHIPS.

IN HONOR OF

PROFESSOR EDWARD T. FRISTOE, WILLIAM S. THOMPSON,
PHAR. D., AND JOHN A. MILBURN, PHAR. D.

THE NATIONAL COLLEGE OF PHARMACY, desiring to honor the memories of the late Prof. EDWARD T. FRISTOE and Messrs, WILLIAM S. THOMPSON and JOHN A. MILBURN, all of whom while living rendered great and acceptable services to this College, has therefore established three memorial scholarships that are to be known as the Edward T. Fristoe Memorial Scholarship, which is open only to Freshman students ; the William S. Thompson Memorial Scholarship, which is open only to Junior students, and the John A. Milburn Memorial Scholarship, which is open only to Senior students. The scholarships are open to competition under the following conditions:

Students of the Freshman Class who have attended a course of pharmaceutical instruction for the first time and in this College only will be eligible to compete for the Edward T. Fristoe Scholarship.

Students of the Junior Class who have attended and passed a satisfactory examination in the Freshman Class will be eligible to compete for the Wm. S. Thompson Scholarship.

Students who have attended both the Freshman and Junior Classes and passed a satisfactory examination in both will be eligible to compete for the John A. Milburn Scholarship.

The students in the respective classes who at the end of the term pass the required examinations and obtain the highest average in all the branches taught, and who have regularly attended the lectures and performed the amount of laboratory work required during the specified course, will be entitled to receive the scholarship upon a report with recommendation to that effect from the Board of Examiners, and the student to whom the scholarship has been awarded shall receive gratuitously and in their due order the lectures and laboratory instruction appertaining to that scholarship and only for the year that it has been granted.

RECIPIENTS, 1911.

EDWARD T. FRISTOE SCHOLARSHIP.

CLYDE E. SNIDER.

WILLIAM S. THOMPSON SCHOLARSHIP.

LEO C. THYSON.

JOHN A. MILBURN SCHOLARSHIP.

LORING W. BEESON.

TEXT BOOKS.

Remington's Pharmacy, Waggaman's Compendium of Materia Medica; Gray's Elementary Botany; United States Pharmacopœia; Qualitative Chemical Analysis, Prescott and Johnson; Tanner's Poisons; Gage's Principles of Physics, Goodspeed's edition; Remsen's Introduction to Chemistry; Remsen's Introduction to the Compounds of Carbon; Simon's Manual of Chemistry; Scoville's Art of Compounding; Culbreth's Materia Medica and Pharmacology; Greenish, Foods and Drugs.

BOOKS FOR REFERENCE.

Pharmaceutical and Medical Chemistry, Sadtler and Trimble, second edition, volume 1; Roscoe's and Schorlemmer's Chemistry; Taylor's Toxicology; Maisch's Organic Materia Medica; Bastin's Botany; United States and National Dispensatories, Incompatibles in Prescriptions, Ruddiman; Morphology and Histology of Plants, Rusby & Jelffe; Caspari's Practice of Pharmacy; Pharmacology, Tyrode.

EXAMINATIONS.

The annual examination of students will begin on Friday, April 26, 1912, and continue upon such days as the Board of Examiners may determine.

Students intending to present themselves for the annual examination must exhibit to the Dean of the College on or before April 20, 1912, their lecture tickets, with a recommendation and a statement of attendance endorsed on the back thereof by the several members of the faculty. Candidates for graduation are required to deposit with the Dean the Diploma Fee, which will be returned in the event that the candidate is unsuccessful.

CREDITS ON EXAMINATIONS FOR GRADUATION.

Students who fail to pass the final examination in all studies but are successful in two or more of them shall receive credit for those passed, provided they pass the examination in the remaining studies within two years thereafter; otherwise no credits will be allowed.

THE DEGREE.

The Degree conferred by this College is that of *Doctor of Pharmacy*.

QUALIFICATIONS FOR THE DEGREE.

1. He shall have attended three *courses of instruction* in Chemistry, Pharmacy, Analytical Chemistry, Materia Medica, Botany and

Toxicology, and two in Microscopy, *the last of which* must have been in this College; and one course each in Mercantile Pharmacy and Pharmaceutical Jurisprudence.

2. He must have passed a satisfactory examination in *each* of the branches taught.

3. He must be recommended by the Board of Examiners.

FEES.

Matriculation,	-	-	-	-	-	-	-	-	\$ 5 00
Tickets for the full year's course of instruction, Freshman	-	-	-	-	-	-	-	-	75 00
" " " " " " Junior,	-	-	-	-	-	-	-	-	85 00
" " " " " " Senior,	-	-	-	-	-	-	-	-	90 00
Single tickets for Materia Medica, Botany and Toxicology, each,	-	-	-	-	-	-	-	-	15 00
" " " Analytical Chemistry,	-	-	-	-	-	-	-	-	20 00
" " " Chemistry,	-	-	-	-	-	-	-	-	20 00
" " " Practical Pharmacy,	-	-	-	-	-	-	-	-	20 00
" " " Microscopy,	-	-	-	-	-	-	-	-	15 00
" " " Mercantile Pharmacy,	-	-	-	-	-	-	-	-	5 00
" " " Pharmaceutical Jurisprudence,	-	-	-	-	-	-	-	-	5 00

Fee for Diploma, \$10.00.

SINGLE TICKETS.

Students who desire may select one or more of the branches taught, and upon payment of the fee for single tickets, attend the lectures and laboratory work during the time set apart for such study.

Students taking single tickets will not be entitled to take the examinations for the degree conferred by the College.

Students will not be allowed to attend any course of instruction unless they are provided with the proper tickets. Tickets will be accepted for the course of lectures only for which they are issued. Students failing in the examination for graduation will be required oh take out tickets for and attend instructions at least one year in such studies as they fail to receive credit for before they will be givet another examination.

A deposit of \$3.00 to cover loss by breakage, etc., will be required in each of the two Laboratory Courses. A deposit of \$5.00 will be required of special students in the Analytical Laboratory. Any balance remaining to the credit of the student at the end of the school year will be returned if applied for within one year after the end of the term *and all dues to the College have been paid.*

Tickets can be obtained from the Dean of the College, to whom application for information should be made.

H. E. KALUSOWSKI, Dean,

808 I Street, N. W., Washington, D. C.

UNDERGRADUATES.

1911-1912.

Miss GRACE L. ANDERSON, Mo.
CHAS. W. BARKER, Calif.
CLYDE BUCHANAN, N. C.
WM. A. BOYD, D. C.
FORREST P. BARNES, Ohio.
CECIL E. BROCKMAN, Va.
BENJAMIN CLAR, Russia
RALPH V. CHAMBLIN, Va.
C. F. W. DAMMEYER, Md.
HAROLD G. DAY, D. C.
HERBERT A. DALY, D. C.
FRED'K E. DUDLEY, Jr., D. C.
FRANCIS T. DONAHUE, D. C.
MRS. E. T. ELLIOTT, Mich.
F. M. FELLER, Va.
ALBERT F. GORSUCH, D. C.
L. R. GRUBBS, Md.
L. M. HARBAUGH, D. C.
W. K. HENDERSON, Va.
C. W. HENRY, N. C.
S. J. HOHBERGER, D. C.
H. D. HUGHES, D. C.
RAYMOND D. KINSEY, D. C.
E. A. KENNER, D. C.
ALBERT M. KLOCZEWSKI, D. C.
DANIEL G. LUCKETT, D. C.

JOSEPH LEAR, Russia.
F. B. MARSDEN, D. C.
REDMOND MAYO, N. C.
MALCOLM W. MORGAN, Md.
Miss ANNA E. MIX, Md.
H. M. MILLER, B. S., Ky.
DAVID L. MAXWELL, Tenn.
GEO. E. McCANN, Mass.
W. N. NORTON, Wash.
JOSEPH M. NEIL, Pa.
J. B. SCHOMMER, LL. D., Min.
THOS. F. SCHWEINHOUT, D. C.
H. P. SENAY, D. C.
BERT H. SMYSER, Pa.
EDWIN G. SWANN, Md.
MILES T. SHIPMAN, Kans.
CLYDE SNIDER, E. Okl.
LEO. C. THYSON, D. C.
MILTON H. THOMPSON, D. C.
IRVING A. TENNYSON, Va.
JOHN W. UMHAU, D. C.
FRED W. WALKER, Va.
F. LESLIE WIGHT, Va.
HERBERT F. WHITE, Va.
WILLIAM H. WHITTLESEY, Ohio.
LAWRENCE B. WHITLEY, N. C.

CLARENCE H. WILEY, Va.



STUDENTS TAKING SPECIAL COURSES.

1911.

FRANK L. ANKERS, Va.
JOHN T. PEVARE, N. H.
FRED. E. T. OFFUTT, West Va.



GRADUATES.

1911.

HOMER K. BUTLER, Md.
LORING W. BEESON, Iowa.
WILLARD DAY BOWER, Md.
RALPH W. FELLER, Va.
CHARLES B. GASS, Md.

THEODORIC L. GILL, Va.
WILLIAM SIDNEY JONES, Va.
Miss GAIL E. NELSON, B. S., So. Dak.
CARL F. SNIDER, D. C.
CHARLES WHITEBREAD, Wisc.

ALUMNI.

NOTE.—Those having a * before their names have received certificates, and not the degree conferred by the College. Those having a † before their names are deceased.

HONORARY MEMBERS.

†B. F. CRAIG, M. D. OSCAR OLDBERG †RICHARD H. STABLER, M. D. †EDWARD T. FRISTOE, LL. D. J. EDWARD WHITFIELD. †WILLIAM S. THOMPSON. HENRY A. JOHNSTON. FREDERICK A. HOLTON. FRANK P. WELLER	†JOHN A. MILBURN CHAS. BECKER. WM. HALLOCK. †W. G. DUCKETT. †GILES G. C. SIMMS. †RANDOLPH L. ELIOT. WILLIAM F. HILLEBRAND †A. J. SCHAFHIRT.
--	--

CLASS 1873

C. R. DUFOUR, D. C. ALBERT M. READ, Michigan. †F. C. GAITHER, D. C. †C. LeROY SAYRE, New York. A. S. TABER, New York. E. M. TABER, New York HENRY ADAMS, Massachusetts. FRANK PITZER, Virginia. †T. M. COOMES, D. C. †T. G. DeMOLL, D. C. †C. G. DULIN, D. C. C. M. BALL, Virginia. F. McC. CRISWELL, D. C. SAMUEL BOWER, New York. HARRY W. HODGES, Georgia, NELSON HEAD, Virginia THOS. G. McGORK, California. LOUIS KOLIPINSKI, D. C. †M. MUNCASTER, Maryland. CHAS. ROSCOE LUCE, New York. C. P. KEARFOOT, West Virginia. GEORGE E. DOERING, D. C. FRANK C. HENRY, D. C. S. EDGAR MAHAN, Delaware. †JOS. R. WALTON, M. D., England. ALEX. MUNCASTER, Maryland. JOHN C. FIRMIN, Massachusetts. †WM. S. THOMPSON, Jr. D. C. †E. CHESTER STOTT, D. C. †EDWARD R. BIGELOW, Ohio. LEWIS FLEMER, D. C. B. B. OWEN, Maryland. JAS. F. R. APPLEBY, D. C. CHARLES A. BECKER, New York. R. BENNETT, Ireland. G. R. LEE COLE, Virginia. JOS. A. HORIGAN, D. C. R. L. LYNCH, D. C. CONRAD WEISS, D. C. †MAX KOCH, Illinois. FRANCIS St. CLAIR, New York. W. C. NESS, D. C. E. GEO. BECKER, Germany. †CARL L. CLUSS, D. C. HERBERT C. EASTERDAY, Virginia. †O. R. LATHAM, Pennsylvania. E. N. MATTHEWSON, Pennsylvania. J. A. RIGGS, Maryland.	†WM. B. HEISKILL, Pennsylvania. T. P. COLE, D. C. H. E. KALUSOWSKI, D. C. †W. F. SCALA, D. C. J. SNIDER NOEL, Maryland. †D. F. OWENS, Maryland. †T. E. CHIDESTER, Ohio. †JOHN J. STAFFORD, Maryland. I. B. ROBERTSON, New York T. A. T. JUDD, Virginia. WM. E. SHAFFER, D. C. HARRY STANDIFORD, Virginia. †ANDREW F. HOFER, Pennsylvania. EDWIN GLADMON, D. C. GEORGE W. BOYD, D. C. WM. HINRICHS, Germany. HENRY EVANS, England. †LEWIS C. MILBURN, Virginia. EDWARD P. MERTZ, D. C. WM. T. CRISWELL, D. C. JAMES A. WATSON, Virginia. ALBERT E. ACKER, D. C. †CHARLES F. KEIM, Pennsylvania. GEORGE J. LOCHBOEHLER, Missouri. LAIDLER MACKALL, D. C. ALBERT A. LAWRENCE, Ohio. WM. K. MITCHELL, D. C. †ALFRED MOSS, Virginia. E. C. F. A. SCHAEFER, Germany. HERMAN B. WADDY, Virginia. W. G. ALDRIDGE, Virginia. ED. V. CONNER, D. C. F. HAFELFINGER, New York. C. H. FRANZONI, D. C. H. W. SESSFORD, D. C. E. F. SICKENBERGER, Egypt. †W. L. SKINNER, Virginia. SAMUEL T. STOTT, D. C. JOHN E. TONER, D. C. E. W. WHITESIDE, Maryland. †W. E. WOLHAUPTER, D. C.
--	---

- 1888
 +HIRAM J. BANES, Pennsylvania.
 ARTHUR B. BURROWS, New York.
 HUGH M. CLINE, Maryland.
 WM. B. CULVERWELL, D. C.
 V. H. EISENBEISS, D. C.
 +FRANK D. EVANS, D. C.
 WM. H. FREY, D. C.
 CHARLES E. GROSS, D. C.
 CHAS. HAWKINS, England.
 GEO. W. HURLEBAUS, D. C.
 EDWARD A. HELMSEN, D. C.
- EDWARD BOYD, D. C.
 J. S. CLEMENCE, D. C.
 WM. EMORY, Pennsylvania.
 W. W. FISHER, Pennsylvania.
 C. J. GILLETTE, New York.
 S. T. GRIMES, Maryland.
- 1889
 W. P. HEREST, Pennsylvania.
 J. S. HIGDON, Maryland.
 ROBERT L. WREN, Virginia.
 M. H. MENKE, D. C.
 C. J. ORDING, Illinois.
 GEO. B. WEISS, D. C.
- 1890
 +ISAAC ALLEN, Delaware.
 ROBERT C. DICKINSON, Pennsylvania.
 MATTHEW B. DONNELLY, D. C.
 CHARLES EARL, Massachusetts.
 WALTER S. FERRIS, New York.
 MONTE GRIFFITH, Virginia.
 CALVIN B. HEIZER, Virginia.
 R. VERNON HOUSTON, D. C.
 +PERCY G. McCOMAS, Maryland.
- EDWARD F. ALBERT, Pennsylvania.
 CHARLES E. BALDWIN, Iowa.
 ALFRED T. BRONAUGH, Virginia.
 CHAS. B. CAMPBELL, Pennsylvania.
 +J. EDWARD CARROLL, D. C.
 ALBERT J. COX, Virginia.
 FRANCIS C. HAINES, D. C.
 JAMES W. HARPER, Virginia.
 HENRY S. HERR, Illinois.
 HARRY D. HUTTON, D. C.
- 1891
 ALBERT B. HYATT, Maryland.
 JOHN W. JENNINGS, Virginia.
 FRANK V. JOHNSON, D. C.
 CHARLES MacGREGOR, Virginia.
 +WILLIAM G. ROE, Pennsylvania.
 +WILLIAM J. RYDER, New York.
 CHARLES SCHERER, D. C.
 ALFRED H. WELLS, Maryland.
 GILPIN WILLSON, Maryland.
 +ROBERT C. WILLIAMS, Virginia.
- 1892
 SMITH C. PEDIGO, Texas.
 ADOLPHUS E. POWELL, Virginia.
 TOM N. PHILLIPS, Colorado.
 Miss JENNIE M. REIGART, Colorado
 FELIX A. VAN REUTH, D. C.
 HARRY A. YATES, Virginia.
- 1893
 GEORGE J. GEIGER, D. C.
 +WILLIAM E. HALLORAN, D. C.
 DANIEL D. MULCAHY, D. C.
 *WILLIAM S. UDALL, Vermont.
 FRANK C. WILSON, Ohio.
- 1894
 ALFRED H. KEIM, Virginia.
 JOHN A. KOCH, Illinois.
 W. KERFOOT LUPTON, D. C.
 QUENTIN MACKALL, D. C.
 HERBERT D. MEEK, Pennsylvania.
 CHARLES C. MUSSINA, Pennsylvania.
 GUY M. NEELY, Pennsylvania.
 EUGENE R. NICHOLS, D. C.
 *TERRY OLESON, Minnesota.
 LOUIS RUBENSTEIN, D. C.
 THOMAS J. SHERIDAN, D. C.
 MORGAN L. STEELE, N. Y.
 ROBERT A. VEITCH, Maryland.
- 1895
 FRED B. HASKINS, D. C.
 JAMES A. JENNINGS, Virginia.
- 1896
 +CHARLES W. LITTLE, Kansas.
 JOHN T. MURPHY, Wisconsin.
 +ARTHUR L. ORRISON, Virginia.
 HARRY M. PRICE, D. C.
 *JENNIE T. RUGG, Massachusetts.
 +CHARLES G. SANDERS, Mo.
 DANIEL F. SLATTERY, D. C.
 ODEN R. SUDLER, D. C.
 M. R. WOOLDRIDGE, D. C.
 C. HOWLE YOUNG, D. C.
- +JAMES T. ARWINE, Indiana.
 EDWARD J. BASTABLE, D. C.
 ALBERT N. CONNER, Virginia.
 WM. M. JOHNSON, D. C.
 W. P. M. KING, D. C.
 GEORGE T. MANKIN, Virginia.
- WYMOND H. BRADBURY, New Jersey.
 HENRY BUDENBOM, Indiana.
 *CHARLES L. EBAUGH, Mo.
 ARTHUR CASE FITCH, New York.
 *HENRY R. GARLAND, Virginia.
- PHILLIP J. AFFLECK, Jr., Virginia.
 J. THOMAS BAILEY, Maryland.
 MELVILLE L. BRADFORD, D. C.
 +CHARLES H. BLUMER, D. C.
 JULIUS S. BUYNITZKY, D. C.
 NEWTON EDMONDS, D. C.
 ERNEST T. FEARON, Pennsylvania.
 THORNTON B. FISHER, D. C.
 WALTER R. HILL, D. C.
 HARRY T. L. HOYLE, D. C.
 BERTRAM A. JOHNSON, D. C.
 V. MASON JOHNSON, Virginia.
 CHRISTIAN L. KRAUS, D. C.
 *HARRY B. KREBS, Pennsylvania.
- *MAX GEORGII, Minnesota.
 CHARLES W. HOGAN, Maryland.
 JOHN LEADBEATER, Virginia.
- *HUGH M. ADAMS, Pennsylvania.
 HARRY P. BAKER, D. C.
 +LON BUDENBOM, Indiana.
 MARION E. BULLOCK, Kansas.
 LANGDON S. DAY, Maryland.
 +HARRY T. DODGE, D. C.
 VICTOR H. ESCH, D. C.
 W. CALHOUN FURR, Virginia.
 RICHARD GIBSON, Virginia.
 *HARRY L. GOULD, D. C.
 GEORGE LATTENER, D. C.

- 1897
- * ALEXANDER LEWIS BOGAN, D. C.
 - HARVEY D. BOWDEN, New Jersey.
 - * CLIFTON POWER CLARK, Miss.
 - CLIFFORD S. DUNCAN, Illinois.
 - MARTIN S. FEALY, D. C.
 - JOSHUA L. GATCHEL, D. C.
 - JOHN MURRAY HACKETT, New York.
 - WILLIAM H. HOUGH, Kansas.
 - JOHN W. HOUSTON, Pennsylvania.
 - WILLIAM P. KENEALY, D. C.
 - † ARTHUR H. F. LUERSEN, D. C.
- 1898
- † EDWARD ALVA DUCKETT, D. C.
 - BARRON R. FRANKLIN, Va.
 - CHARLES I. GRIFFITH, D. C.
 - * KIRK HOLMES, N. Y.
 - PRESTON C. KING, D. C.
 - * LUCRETIA B. LACY, Ill.
 - CHARLES A. McAVOY, D. C.
 - E. L. MASON, Va.
- 1899
- LOUIS FRANCIS BRADLEY, Va.
 - CHARLES J. FUHRMANN, Mich.
 - JOHN S. GALLAGHER, D. C.
 - PAUL L. JOACHIM, D. C.
 - GEORGE DEXTER KEHOE, Ky
 - FRANK B. KETCHUM, Mich.
- 1900
- CHARLES GARRELS, Ill.
 - FLORENCE VIRGINIA HOSKINS, D. C.
 - J. WILLARD McCHESNEY, D. C.
 - JOHN M. MINICK, Pa.
 - ERNEST PARSONS, D. C.
 - † WILLIAM A. WOODFIN, Tex.
- 1901
- ALEXANDER SHIRAS DAGGETT, N. Y.
 - F. PERKINS DEWEY, Jr., Tenn.
 - PETER JOSEPH DUNCAN, Conn.
 - JOSIAH H. HOLLAND, D. C.
 - * CHARLES E. HOUGHTON, Mass.
 - ADAM KEMBLE, Pa.
 - WILLIAM T. KERFOOT, Jr., Va
 - * FRANK A. TUCK, Va.
- 1902
- HOWARD M. BRADBURY, Pa.
 - JOHN F. BUTLER, D. C.
 - MANLEY J. CLARK, N. Y.
 - FRANK R. DAVIS, Ky.
- 1903
- W. EVEREST BOYER, Md.
 - WILLIAM A. BRIGGS, Ill.
 - † HARRY A. CANDEE, Ill.
 - LOUIS B. CASTELL, D. C.
 - THOMAS E. GIBB, Scotland.
 - WILLIAM P. HABEL, Pa.
 - HARRY E. HARVEY, Ohio.
- 1904
- J. FOSTER ALLISON, Ill.
 - ISADORE B. COHEN, D. C.
 - FREDRICK Y. DONN, D. C.
 - WILBUR S. HAUER, D. C.
 - CLAUDE E. KOSS, D. C.
- 1905
- JOSEPH F. ARTH, D. C.
 - MATTHEW J. BEISTLE, Mich.
 - FRED B. CAMPBELL, Va.
 - ALBERT P. CLARK, Pa.
 - † ROBERT G. HARVEY, Ohio.
 - SYLVERN LAUPHEIMER, Va
- 1906
- † FREDERICK C. BENNETT, England.
 - ADDIE P. S. CRISWELL, D. C.
 - LOUIS V. DIETER, Md.
 - ALICE WINANS DOWNEY, Ohio.
 - ISADORA GEOGHEGAN, D. C.
 - MILTON L. GOLDSMITH, D. C.
 - JOHN W. GRADY, D. C.
 - BERNARD S. JUDD, D. C.
 - BERNARD B. LARRICK, Va.
- * CLARKE J. MORRISON, Indiana.
 - W. ROBERT PERKINS, D. C.
 - JOHN MATTHEW PULLIAM, Virginia
 - * ALEX. HANSON QUARLES, Georgia.
 - HENRY WRIGHTSON SMITH, Virginia
 - HARRY C. SNYDER, D. C.
 - JOHN P. THOMPSON, D. C.
 - FRANK BROUGHTON TIPTON, D. C
 - CHARLES S. WALTER, D. C.
 - OTTO F. WELLENREITER, Illinois.
- SALVADOR D. MOORE, D. C.
- ALPHONSUS A. O'DONOGHUE, Me
- FRANK C. PURDUM, Md.
- FRANK R. RICHARDSON, Ohio.
- ISAAC SCOTT, D. C.
- ANDREW J. SHERIDAN, D. C
- LLOYD T. TAYLOE, Va.
- TIMOTHY T. LANE, D. C.
- FRED A. MALTBY, D. C.
- ARTHUR W. PARKER, D. C.
- * FRANK J. PHELPS, Pa.
- SAMUEL A. A. RICHARDSON, Ohio
- HERBERT W. POOLE, Va.
- SAMUEL P. RICKARDS, Pa.
- † DEHAVEN SHARP, D. C.
- S. MASON WAGNER, W. Va.
- H. MCCOY WALTERS, Ill.
- J. ARTHUR KLINGER, Pa.
- JOHN KRAUS, D. C.
- W. FENWICK MATTINGLY, Md.
- HELEN M. PROCTOR, Vt.
- FREDERICK REPETTI, D. C.
- JESSE A. SIMPSON, Md.
- THOMAS STRETTON, England.
- PRESTON C. DAY, Md.
- MALCOLM G. GIBBS, Tenn.
- H. K. KIRBY, Va.
- LAWRENCE P. NOLAN, Md.
- FRANCES P. HUTCHINSON, D. C.
- LEWIS H. LAMB, D. C.
- * HENRY C. LEHMANN, D. C.
- W. J. McNAMEE, Pa.
- WILLIAM A. MESS, Ind.
- BENJ. F. SHOWALTER, Va.
- WILLIAM F. WORK, D. C.
- JOHN J. McLOONE, Ireland.
- CHARLES C. READ, Pa.
- JOSEPH C. WILLIAMS, Md.
- JAMES B. WINGATE, Md.
- ALFRED WOLLBERG, Pa.
- JOHN C. PEACOCK, Md.
- SAID T. SAMAHA, Syria.
- CHARLES C. SMALL, Ohio.
- ELMER L. SPITTLE, Va.
- ELIJAH W. TITUS, Va.
- GEORGE S. WEBB, Minn.
- FRANK T. LINTON, Md.
- HARRY S. McAULEY, D. C.
- DANIEL J. MATTINGLY, Md.
- ROBERT R. MISKIMON, Del.
- CYRUS W. NELSON, Iowa.
- AGNES M. NORDEMAN, Ills.
- HELEN HAZEL NORDEMAN, Ills
- NELLIE G. O'DONNELL, D. C.
- HELEN A. SUDLER, Ills.

1907

HERMAN H. COLBY, Germany.
WILLIAM H. CANTWELL, D. C.
BERT V. CUPPERNELL, Ills.
JOHN R. JACOBS, New York.
JOHN T. KEISTER, Va.
CLAUDE J. KEM, Colo.
LOUIS LAUBINGER, Germany.

ROBERT F. TROXLER, Kentucky.

HAROLD H. LANTZ, Va.
ROBERT EMMETT MADIGAN, D. C.
STELLA C. NELSON, Oklahoma.
ANDREW J. O'NEILL, D. C.
BENNO R. PREUSS, Texas.
LOUIS SACHS, Germany.
W. BURTON SPIRE, New York.

1908

W. R. BOYER, Md.
CARROLL G. DEMING, D. C.
HENRY B. FLOYD, Arkansas.
MORRIS A. POZEN, Russia.

ROBERT B. SPENCER, N. C.
ERNEST H. STEELE, Va.
TAYLOR O. TIMBERLAKE, Va.
WILLIAM D. THORN, D. C.

1909

RAY T. BAILEY, D. C.
T. QUINN JONES, Md.
EDWARD V. PAYNE, Va.

CHAUNCEY C. REESE, Md.
IRENE NELLIE RICHARDSON, D. C.

1910

GEO. W. F. BOYD, D. C.
ELIAS ELVOVE, B.S., M S , Pa.
J. D. A. HOGAN, D. C.
RALPH A. JUDD, D. C.
ALBERT W. KENNER, Mass.
J. HAROLD MORGAN, Md.
FRANR W. MILBURN, Va.

JAMES I. NOLAN, Ills.
PAUL E. PLUNKETT, Md.
DAVID B. PETERS, Va.
NAOMI E. RICHARDSON, D. C.
MARY H. RICHARDSON, Ohio.
JULIA H STROBEL, D. C.
MELVILLE B. TEWKSBURY, Kan.

DOUGLAS TSCHIFFELY, Md.

1911

HOMER K. BUTLER, Md.
EORING W. BEESON, Iowa.
WILLARD DAY BOYER, Md.
THEODORICL. GILL, Va.
CHARLES B. GASS, Md

CHARLES W. HENDERSON, Va.
Mrs. GAIL E. NELSON, B. S., So. Dak.
WILLIAM SIDNEY JONES, Va.
CARL F. SNYDER, D. C.
CHARLES WHITEBREAD, Wisc.

UNIVERSITY OF ILLINOIS-URBANA



3 0112 110184550

Barrett Printing Company
Washington, D. C.